

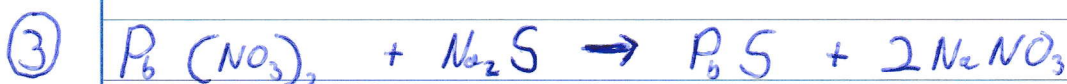
ANTWOORDEN PROEFTOETS ANORGANISCH

$$\begin{aligned} \textcircled{1} a \quad & \text{Ca} = 40,08 \times 3 = 120,24 \\ & \text{P} = 30,97 \times 2 = 61,94 \\ & \text{O} = 15,99 \times 8 = 127,92 + \\ & \qquad \qquad \qquad 310,07 \text{ g/mol} \end{aligned}$$

$$\begin{aligned} b \quad & \text{C} = 12,01 \times 4 = 48,04 \\ & \text{H} = 1,008 \times 8 = 8,06 + \\ & \qquad \qquad \qquad 92,13 \text{ g/mol} \end{aligned}$$

$$\begin{aligned} \textcircled{2} a \quad & \text{C} = 12,01 \times 1 = 12,01 \qquad \frac{12,01}{43,99} \times 100\% = 27,30 \\ & \text{O} = 15,99 \times 2 = 31,98 + \\ & \qquad \qquad \qquad 43,99 \text{ g/mol} \end{aligned}$$

$$\begin{aligned} b \quad & \text{C} = 12,01 \times 2 = 24,02 \qquad \frac{24,02}{46,06} \times 100\% = 52,15\% \\ & \text{H} = 1,008 \times 6 = 6,05 \qquad 46,06 \\ & \text{O} = 15,99 \times 1 = 15,99 + \\ & \qquad \qquad \qquad 46,06 \end{aligned}$$



$$n = \frac{m}{M} \rightarrow \frac{30}{78,05} = 0,38 \text{ mol Na}_2\text{S}$$

$$\text{Reactie} = 1:1 \rightarrow 0,38 \text{ mol P}_2\text{S}$$

$$m = n \cdot M \rightarrow 0,38 \times 239,24 = 91,57 \text{ g P}_2\text{S}$$



$$n = c \cdot V \rightarrow 0,0986 \times 17,65 = 1,74 \text{ mmol OH}^-$$

$$\text{Reactie} = 1:1 \rightarrow 1,74 \text{ mmol CH}_3\text{COOH}$$

$$m = n \cdot M \rightarrow 1,74 \times 60,05 = 104,50 \text{ mg CH}_3\text{COOH}$$

$$\frac{104,5}{2578} \times 100\% = 4,05\% \text{ (m/m) CH}_3\text{COOH}$$

$$2578$$

⑤



$$n = \frac{m}{M} \rightarrow \frac{151,4}{106,0} = 1,43 \text{ mmol Na}_2\text{CO}_3$$

$$\text{Reactie} = 1:2 \rightarrow 2,86 \text{ mmol HCl}$$

$$c = \frac{n}{V} \rightarrow \frac{2,86}{28,13} = 0,1016 \text{ mmol/ml HCl}$$

ml/L
M